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Photoelectron Spectra, MNDO Calculations and Electronic Structure of Some Saturated Steroids

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Photoelectron (PE) spectra in connection with semiempirical MNDO SCF MO calculations (assuming validity of Koopmans' theorem) and empirical arguments (*i. e.* composite molecule method) are used to derive (valence) electron structure of 5 α -androsterane (1), 5 α -androstan-3-one (2), 5 α -androstan-11-one (3), 5 α -androstan-17-one (4), 5 α -androsterane-3,17-dione (5) and 5 α -androsterane-3,11,17-trione (6).

INTRODUCTION

In the past years we have been interested in PE spectra of steroids.¹⁻³ Steroids, a group of alicyclic compounds containing a perhydrogenated cyclopentanophenanthrene («sterane» or «gonane») as a basic structural characteristic, play a variety of biological roles in humans. They are components of bile, brain, spinal tissue and some of them are sexual hormones. In a number of cases, the mechanism of hormone action involves interaction with specific receptors in cellular plasma membranes and subsequent generation of intracellular secondary messengers (such as cyclic-AMP, cyclic-GMP, diacylglycerol, calcium ion *etc.*). In other words, the information that is inherent in an extracellular hormone is transferred and amplified by an intracellular mediator for that hormone. Steroid hormones utilize a different mechanism of information transfer.⁴⁻⁶ Essential steps in induction of a biological response are the binding of a hormone to a receptor protein present in a target cell and the interaction of the receptor-protein complex with nuclear chromatin. The precise details of this interaction and the role played by the steroid in this process are still unknown. But molecular structure (*i. e.* composition, configuration, conformation) and electronic properties (*i. e.* energy, electron distribution) undoubtedly have a direct bearing upon receptor affinity and directly or indirectly influence receptor activation, transport and nuclear interaction. The structural factor is well documented. From the comparison of the three-dimensional structures and binding affinities of the steroid hormones, their agonists and antagonists, it is possible to deduce structural characteristics compatible with binding to the protein, while from the know-

ledge of the amino-acid sequence and predicted secondary structure of the protein it is possible to identify putative binding sites by modeling and/or graphic techniques. On the other hand, our knowledge about the electronic structure of steroids, both by quantum chemistry and PE spectroscopy, is very sparse. Thus, elucidation of (valence) electronic structure of steroids is essential for the better understanding of their biochemistry.

Mainly because of their well defined stereochemistry, steroids have been widely used in the study of cooperative long-range substituent effects (through space *vs.* through bond) at distances fixed by an inert spacer — steroid framework. But recent studies on intramolecular electron-transfer (ET) processes in molecular systems in which donor and acceptor groups are separated by 5 α -androstane⁷ (or by some other saturated polycyclic hydrocarbon⁸) question that long-standing hypothesis about the inertness of the basic steroid skeleton. The theoretical and experimental evidences suggest that the coupling between donor and acceptor in these systems is provided by the interaction with the frontier orbitals of the intervening molecular fragments. It follows that the »ribbon-like« orbitals of saturated polycyclic hydrocarbons can »conduct« electrons over ten or even more single CC bonds, although more experimental evidence is needed. In this context, the informations provided by PE spectroscopy and/or quantum chemical calculations are of crucial importance.

In this work, we investigate the long-range interactions of carbonyl groups located at the biologically important 3- and/or 11- and/or 17- positions of steroid framework. The information about the type and extent of interactions between them, as well as on MO symmetries, is provided by the MNDO calculations. This method gives a qualitatively correct description of the valence electronic structure of these molecules.

EXPERIMENTAL AND CALCULATIONS

The Hel PE spectra of compounds 1, 2 and 4–6 were recorded on a Vacuum Generators UV-G3 spectrometer⁹ at 80, 130, 140, 170 and 180 °C, respectively, using MeI and/or Ar and/or Xe as calibrants in admixture with the compound. Compounds were of commercial origin and were used without further purification.

Standard semiempirical MNDO SCF MO calculations were performed for all compounds using the QCPE 464 (IBM MOPAC) program¹⁰ adapted for the Siemens 7580S (turbo-version) at the Computing center of the University of Düsseldorf. Full geometry optimization was obtained for 11-one and 3,11,17-trione, while for the other measured compounds single geometry calculations were performed. Input data were chosen on the basis of crystallographic data¹¹ and standard values of bond lengths and angles ($r_{CC} = 154$ pm, $r_{CH} = 108$ pm, $r_{CO} = 122.5$ pm, $\angle CCC = 109.5^\circ$ and 120.0° , $\angle CCH = 109.5^\circ$, $\angle COC = 120.0^\circ$).

RESULTS AND DISCUSSION

The MNDO energies and assignment for the oxygen lone pair of compounds 1–6 are compared to the corresponding PE ionization energies (Table I). This comparison is being made on the assumption that the overall shape of the observed PE spectrum could be obtained, after appropriate scaling, by replacing each orbital eigenvalue by a Gaussian of width 0.250–0.300 eV.¹² The non-zero (>0.001) eigenvectors for the six highest occupied MO's of the same molecules are given in Table II.

TABLE I

Canonical orbital energies, ϵ_i^{MND0}/eV , for the thirty highest occupied levels in 5 α -androstane derivatives 1–6 and measured ionization energies (I_p/eV)

	1		2		3		4		5		6	
	ϵ^{MND0}	I_p	ϵ^{MND0}	I_p	ϵ^{MND0}	ϵ^{MND0}	I_p	ϵ^{MND0}	I_p	ϵ^{MND0}	I_p	
HOMO	-10.975	9.21	-10.532	8.92 n_0	-10.367	n_0	-10.376	8.75 n_0	-10.453	8.94 n_0	-10.474	8.90 n_0
2nd	-11.127	9.49	-11.191	9.41	-11.209		-11.339	9.35	-10.601	9.06 n_0	-10.618	9.09 n_0
3rd	-11.515		-11.447		-11.313		-11.548	9.61	-11.511	9.82	-11.059	9.50 n_0
4th	-11.552	10.32	-11.698		-11.698		-11.733		-11.721	10.18	-11.919	
5th	-11.756		-11.908	10.38	-11.819		-11.791	10.56	-11.906		-12.085	10.28
6th	-11.819		-12.034		-11.948		-11.950		-12.188	10.75	-12.243	
7th	-12.006		-12.060		-12.243		-12.210		-12.233		-12.381	10.48
8th	-12.143		-12.196		-12.374		-12.241		-12.429		-12.472	
9th	-12.187	11.94	-12.346		-12.449		-12.410		-12.718	11.09	-12.813	10.94
10th	-12.336		-12.470	11.27	-12.552		-12.551		-12.787		-12.905	
11th	-12.417		-12.632		-12.718		-12.603	11.95	-12.943		-12.983	
12th	-12.529		-12.724		-12.766		-12.694		-12.989		-13.151	12.38
13th	-12.604		-12.797		-12.801		-12.797		-13.031	12.33	-13.167	
14th	-12.691	12.77	-12.881	12.02	-12.877		-12.876		-13.139		-13.278	
15th	-12.709		-12.993		-12.967		-13.003		-13.229		-13.415	12.64
16th	-12.875		-13.176	12.83	-13.090		-13.136	12.72	-13.447		-13.493	
17th	-12.962		-13.410		-13.349		-13.203		-13.558		-13.659	
18th	-13.260	13.57	-13.565		-13.404		-13.489		-13.711	13.34	-13.757	
19th	-13.469		-13.671	13.66	-13.599		-13.753		-13.784		-13.900	13.77
20th	-13.562		-13.711		-13.825		-13.930	13.47	-13.910		-14.009	
21st	-13.748	14.65	-13.855		-13.987		-14.033		-14.007	13.89	-14.255	
22nd	-13.861		-13.942	14.80	-14.062		-14.089		-14.169		-14.357	14.82
23rd	-13.984		-14.045		-14.184		-14.187		-14.356		-14.449	
24th	-14.124		-14.306		-14.209		-14.395	14.0	-14.399		-14.576	
25th	-14.347	15.67	-14.395	15.37	-14.313		-14.598		14.565	15.02	-14.728	15.77
26th	-14.396		-14.627		-14.615		-14.860		-14.607		-14.758	
27th	-14.442		-14.728	15.82	-14.652		-14.965	15.60	-14.806	16.00	-15.074	
28th	-14.944		-15.135	17.14	-14.932		-15.151		-15.299		-15.444	
29th	-15.111	17.02	-15.316		-15.202		-15.335	17.02	-15.377	17.34	-15.521	17.29
30th	-15.141		-15.474	17.62	-15.397		-15.610		-15.558		-15.581	

The highest occupied »ribbon-like« MO's are considered to intervene in long-range effects which are taking part in large polycyclic hydrocarbons. Being formed from C2p AO's aligned along molecular axes, HOMO and SHOMO of 5 α -androstane (Figure 1) have the highest electron density on two inner rings carbon atoms. According to the alignment of the constituent C2p AO's they can be classified as transversal (T-type; HOMO) and longitudinal (L-type; SHOMO) orbitals. The characteristic feature of these orbitals is electron density spread over the whole system, therefore being described as a mixture of σ_{CC} and π_{CH_2} orbitals. The pseudo π_{CH_2} nature of methylene electrons is shown to be essential for transmission of electronic effects in substituted polycyclic hydrocarbons.¹³ High electron density in the middle of the molecular frame would have significant influence on C2p AO's interactions if we were to alter the original structure of B and/or C rings (*i. e.*, by interpolation of double bonds^{3,14}).

Substitution of methylene group by keto group brings about notable changes in the electronic structure of the parent molecule. There is a pro-

Table II (contd.)

sC ₁₅	-0.018	-0.029	0.111	0.022	0.029	-0.033	0.003	0.020	-0.038	0.010	-0.005	-0.042	sC ₁₅	0.002	-0.020	-0.010	0.006	-0.026	0.013
PaC ₁₅	-0.062	0.139	0.008	0.139	0.008	0.098	0.011	0.090	-0.162	0.078	0.164	0.075	PaC ₁₅	-0.022	-0.096	0.015	0.055	-0.127	0.089
PaC ₁₅	0.101	0.088	-0.020	-0.154	0.003	0.016	-0.013	-0.126	0.029	-0.012	-0.126	0.029	PaC ₁₅	0.018	0.004	0.046	-0.058	0.206	0.062
PaC ₁₅	-0.053	-0.060	0.090	0.045	0.162	-0.083	0.006	0.057	-0.081	0.094	-0.080	-0.184	PaC ₁₅	-0.024	-0.028	0.001	-0.029	-0.017	0.175
sC ₁₆	-0.023	0.008	-0.006	0.019	0.008	-0.036	0.006	0.025	0.006	-0.011	-0.028	-0.048	sC ₁₆	0.005	0.010	-0.028	0.001	-0.029	-0.014
PaC ₁₆	-0.021	0.061	-0.035	-0.031	-0.044	-0.162	-0.004	0.011	0.077	-0.053	-0.108	-0.178	PaC ₁₆	0.012	0.064	0.023	0.034	0.005	-0.062
PaC ₁₆	0.003	0.052	0.050	-0.028	0.004	0.187	0.000	0.082	0.082	0.072	0.185	0.185	PaC ₁₆	-0.044	0.007	0.002	0.024	0.024	0.072
PaC ₁₆	-0.023	0.029	-0.019	0.037	0.037	0.037	0.001	-0.003	-0.006	-0.011	0.090	0.073	PaC ₁₆	0.007	0.000	0.002	-0.041	-0.066	-0.066
sC ₁₇	-0.026	0.038	0.011	0.007	0.019	0.037	-0.006	0.029	0.046	0.013	0.016	0.020	sC ₁₇	-0.020	0.070	0.010	0.010	0.010	0.008
PaC ₁₇	-0.151	0.139	0.078	0.110	-0.135	0.303	-0.006	0.184	0.177	0.090	0.184	0.223	PaC ₁₇	-0.122	0.054	-0.062	0.053	0.019	-0.160
PaC ₁₇	0.025	-0.031	-0.072	-0.006	0.039	-0.201	-0.031	-0.031	-0.054	-0.094	-0.059	-0.179	PaC ₁₇	0.061	0.000	0.006	0.024	0.006	0.062
PaC ₁₇	0.011	0.004	-0.021	-0.032	-0.024	-0.053	-0.011	-0.007	-0.039	-0.019	-0.034	-0.034	PaC ₁₇	0.018	-0.024	0.004	-0.028	0.002	-0.034
sC ₁₈	0.007	-0.005	0.027	-0.008	-0.022	-0.031	0.056	0.056	0.048	0.077	0.104	0.122	sC ₁₈	0.016	0.003	0.004	-0.024	0.006	-0.039
PaC ₁₈	-0.046	0.035	0.072	0.055	0.143	0.156	0.002	0.041	-0.048	-0.096	0.041	0.097	PaC ₁₈	-0.011	0.005	-0.031	0.033	-0.064	0.003
PaC ₁₈	-0.035	-0.025	-0.064	0.076	-0.043	-0.108	-0.001	-0.071	-0.009	-0.205	-0.089	0.049	PaC ₁₈	0.009	-0.009	0.000	0.001	-0.064	-0.109
PaC ₁₈	0.051	-0.009	-0.148	-0.063	-0.147	-0.099	0.001	-0.004	-0.006	-0.026	-0.006	-0.005	PaC ₁₈	0.018	0.007	0.049	-0.053	0.069	-0.216
sC ₁₉	0.003	-0.005	0.038	0.004	0.034	0.020	-0.008	0.038	-0.078	0.006	0.006	0.005	sC ₁₉	0.003	0.003	-0.040	-0.022	0.009	0.069
PaC ₁₉	-0.043	0.070	0.028	-0.043	-0.062	-0.046	0.019	0.026	-0.038	-0.073	-0.099	0.047	PaC ₁₉	-0.028	0.061	-0.067	-0.040	-0.022	0.052
PaC ₁₉	-0.043	-0.048	-0.101	-0.072	0.078	0.057	-0.008	-0.026	-0.038	-0.170	-0.073	-0.099	PaC ₁₉	-0.041	-0.062	-0.112	-0.122	-0.136	0.012
PaC ₁₉	0.029	0.007	0.227	0.061	0.159	0.102	0.018	0.009	-0.026	-0.035	0.001	0.003	PaC ₁₉	0.004	0.050	0.019	-0.215	0.005	0.014
sH(C ₂₀)	0.020	-0.005	-0.111	0.031	0.148	0.099	0.025	-0.002	-0.035	0.001	0.003	0.004	sO ₁	-0.001	0.004	-0.003	-0.003	0.007	0.003
sH(C ₂₀)	0.018	-0.015	-0.067	0.028	0.151	0.111	0.025	0.009	0.026	0.014	-0.111	0.128	PaO ₁	-0.612	0.019	0.232	0.111	0.115	0.080
sH(C ₂₀)	0.009	-0.005	-0.083	0.015	0.166	0.117	0.011	0.065	0.074	-0.156	0.118	0.004	PaO ₁	-0.242	-0.026	0.115	0.028	0.008	-0.010
sH(C ₂₀)	0.020	0.030	-0.083	0.023	0.183	0.152	0.003	-0.077	-0.028	-0.004	-0.118	-0.004	PaO ₁	0.104	-0.072	-0.123	0.022	-0.102	0.089
sH(C ₂₀)	0.008	-0.025	-0.154	0.019	0.219	0.170	0.018	0.026	0.074	-0.056	0.012	0.033	sH(C ₂₀)	0.005	0.022	0.106	-0.159	-0.114	0.101
sH(C ₂₀)	0.008	0.010	-0.112	0.034	0.096	0.076	0.015	-0.026	0.009	-0.022	0.063	-0.047	sH(C ₂₀)	0.003	0.022	0.068	-0.166	-0.099	0.142
sH(C ₂₀)	0.005	0.019	-0.141	0.055	-0.035	0.021	0.006	-0.081	0.029	-0.084	0.094	0.094	sH(C ₂₀)	0.007	0.016	0.044	-0.181	-0.086	0.145
sH(C ₂₀)	0.042	0.007	-0.279	0.054	-0.159	-0.004	0.038	-0.042	-0.093	-0.016	0.035	0.035	sH(C ₂₀)	0.007	-0.016	0.088	-0.185	-0.100	0.113
sH(C ₂₀)	0.015	-0.017	-0.303	0.031	-0.028	-0.002	0.038	-0.032	-0.029	-0.016	0.016	0.045	sH(C ₂₀)	0.003	-0.001	0.114	-0.248	-0.117	0.050
sH(C ₂₀)	-0.059	-0.054	-0.200	-0.009	-0.021	-0.092	0.000	-0.035	0.035	0.115	-0.040	0.045	sH(C ₂₀)	-0.005	0.027	0.083	-0.133	-0.074	-0.053
sH(C ₂₀)	0.023	-0.032	-0.133	-0.032	-0.063	-0.166	0.011	-0.039	-0.090	0.020	0.019	0.045	sH(C ₂₀)	0.014	0.055	0.080	-0.068	-0.065	-0.171
sH(C ₂₀)	0.051	0.003	-0.210	0.013	-0.239	-0.007	0.003	-0.148	0.086	0.024	0.024	0.064	sH(C ₂₀)	-0.023	0.037	0.117	-0.071	-0.059	-0.111
sH(C ₂₀)	0.031	0.011	-0.083	0.005	-0.153	0.082	0.004	-0.295	0.134	0.064	0.064	0.064	sH(C ₂₀)	0.023	0.037	0.065	-0.032	0.155	0.067
sH(C ₂₀)	-0.015	0.006	0.026	0.041	-0.082	0.197	0.003	-0.074	-0.279	0.038	-0.004	-0.004	sH(C ₂₀)	0.005	0.018	0.079	-0.053	0.003	-0.338
sH(C ₂₀)	0.061	-0.036	0.023	-0.048	0.066	0.002	-0.022	0.033	-0.053	-0.201	-0.068	-0.004	sH(C ₂₀)	0.005	0.013	0.049	-0.026	0.013	-0.183
sH(C ₂₀)	0.073	0.030	0.031	-0.119	0.096	-0.035	0.007	-0.077	-0.038	-0.092	-0.091	-0.002	sH(C ₂₀)	0.027	-0.012	0.018	0.033	-0.049	0.043
sH(C ₂₀)	0.008	-0.027	0.041	0.009	0.023	0.039	-0.001	-0.052	0.005	-0.248	0.118	0.141	sH(C ₂₀)	-0.027	0.027	0.030	-0.022	0.065	0.043
sH(C ₂₀)	0.043	-0.017	-0.041	-0.067	-0.008	-0.036	0.004	-0.026	0.025	-0.092	0.118	0.171	sH(C ₂₀)	0.017	0.027	0.021	-0.028	0.143	0.124
sH(C ₂₀)	0.123	0.054	0.049	0.185	-0.094	0.004	0.000	0.181	0.020	0.042	0.134	0.196	sH(C ₂₀)	-0.019	0.034	0.023	0.023	0.007	0.098
sH(C ₂₀)	-0.047	-0.151	0.049	-0.120	-0.134	0.004	0.000	0.038	0.049	-0.017	-0.037	0.037	sH(C ₂₀)	0.005	0.015	0.035	-0.025	0.074	-0.036
sH(C ₂₀)	-0.071	0.125	0.016	-0.087	0.010	-0.131	0.002	-0.070	-0.038	0.049	0.041	-0.143	sH(C ₂₀)	0.066	0.091	0.131	0.238	-0.042	-0.032
sH(C ₂₀)	0.120	0.025	0.064	0.190	-0.094	-0.087	0.002	-0.096	0.040	0.040	0.050	0.050	sH(C ₂₀)	0.009	-0.205	-0.089	0.056	0.102	-0.096
sH(C ₂₀)	0.186	0.025	0.094	0.049	-0.007	-0.027	0.002	-0.064	-0.021	-0.069	0.018	-0.058	sH(C ₂₀)	0.034	0.130	-0.162	-0.018	0.104	-0.060
sH(C ₂₀)	-0.079	-0.191	0.049	-0.011	-0.051	0.018	0.000	0.069	0.099	0.112	0.063	0.063	sH(C ₂₀)	0.061	0.074	0.114	0.231	-0.041	-0.038
sH(C ₂₀)	0.180	-0.001	0.092	0.065	0.145	0.116	0.000	0.051	-0.056	0.086	0.086	0.086	sH(C ₂₀)	0.080	0.058	0.114	0.116	0.093	0.098
sH(C ₂₀)	0.082	0.001	0.010	-0.121	0.055	0.003	0.000	0.095	0.103	-0.139	0.070	0.070	sH(C ₂₀)	0.076	-0.186	-0.020	-0.030	0.006	0.046
sH(C ₂₀)	-0.058	0.046	0.088	0.060	-0.038	0.192	-0.001	0.200	0.115	-0.030	0.010	0.010	sH(C ₂₀)	0.011	-0.009	-0.024	-0.040	-0.003	-0.004
sH(C ₂₀)	-0.046	0.033	0.025	0.079	0.093	0.086	0.000	0.029	0.018	0.029	0.116	0.031	sH(C ₂₀)	0.027	0.013	0.058	-0.033	0.146	0.068
sH(C ₂₀)	0.076	0.037	0.065	0.107	-0.043	-0.040	0.000	0.063	0.110	0.124	0.137	0.137	sH(C ₂₀)	-0.021	0.001	0.058	-0.049	-0.062	0.088
sH(C ₂₀)	-0.068	0.076	0.039	-0.075	-0.053	-0.041	0.000	-0.060	0.029	0.049	0.134	0.066	sH(C ₂₀)	0.006	-0.009	-0.009	0.016	-0.097	0.024
sH(C ₂₀)	-0.018	-0.090	0.078	-0.055	-0.043	-0.010	0.000	0.052	0.019	0.049	0.134	0.066	sH(C ₂₀)	0.049	0.060	0.057	0.117	-0.025	0.038
sH(C ₂₀)													sH(C ₂₀)	-0.042	0.054	-0.017	0.065	-0.042	-0.061
sH(C ₂₀)													sH(C ₂₀)	-0.002	-0.124	-0.061	0.047	0.057	0.018

Table II (contd.)

Stc-Androstane-3,11,17-trione		Stc-Androstane-3,11,17-trione		Stc-Androstane-3,11,17-trione		Stc-Androstane-3,11,17-trione		Stc-Androstane-3,11,17-trione		Stc-Androstane-3,11,17-trione		Stc-Androstane-3,11,17-trione		Stc-Androstane-3,11,17-trione	
$E_{\beta}^{(eV)}$		$E_{\beta}^{(eV)}$		$E_{\beta}^{(eV)}$		$E_{\beta}^{(eV)}$		$E_{\beta}^{(eV)}$		$E_{\beta}^{(eV)}$		$E_{\beta}^{(eV)}$		$E_{\beta}^{(eV)}$	
sC ₁	0.005	0.023	0.002	0.002	0.002	0.002	0.002	0.002	0.002	0.002	0.002	0.002	0.002	0.002	0.002
Px ₁ C ₁	0.036	0.049	-0.009	-0.009	0.222	0.075	-0.025	0.010	0.029	0.010	0.029	0.010	0.029	0.010	0.029
Px ₂ C ₁	0.066	0.107	-0.040	-0.040	-0.155	0.082	-0.072	0.025	0.049	0.025	0.049	0.025	0.049	0.025	0.049
Px ₃ C ₁	-0.013	-0.023	0.248	0.248	-0.144	-0.125	-0.125	-0.125	-0.013	-0.023	0.248	0.248	-0.144	-0.125	-0.125
sC ₂	-0.014	-0.080	-0.021	0.005	0.024	-0.003	0.024	-0.003	-0.014	-0.080	-0.021	0.005	0.024	-0.003	0.024
Px ₁ C ₂	-0.064	-0.074	-0.189	-0.189	-0.110	0.018	-0.110	0.018	-0.064	-0.074	-0.189	-0.189	-0.110	0.018	-0.110
Px ₂ C ₂	-0.094	-0.331	-0.051	-0.015	0.013	0.021	-0.051	0.013	-0.094	-0.331	-0.051	-0.015	0.013	0.021	-0.051
Px ₃ C ₂	0.001	0.003	-0.004	-0.117	0.078	0.074	-0.004	-0.117	0.001	0.003	-0.004	-0.117	0.078	0.074	-0.004
sC ₃	-0.005	-0.004	-0.002	-0.009	0.023	-0.009	-0.002	-0.009	-0.005	-0.004	-0.002	-0.009	0.023	-0.009	-0.002
Px ₁ C ₃	0.006	-0.004	-0.008	0.059	0.066	-0.007	-0.008	0.059	0.006	-0.004	-0.008	0.059	0.066	-0.007	-0.008
Px ₂ C ₃	0.072	0.229	0.025	-0.006	-0.002	-0.023	0.025	-0.006	0.072	0.229	0.025	-0.006	-0.002	-0.023	0.025
Px ₃ C ₃	0.029	-0.102	-0.013	0.006	-0.037	-0.040	-0.102	-0.013	0.029	-0.102	-0.013	0.006	-0.037	-0.040	-0.102
sC ₄	0.034	0.078	0.006	0.043	0.037	0.004	0.034	0.078	0.006	0.043	0.037	0.004	0.034	0.078	0.006
Px ₁ C ₄	-0.011	0.137	0.063	-0.124	-0.218	-0.001	-0.011	0.137	0.063	-0.124	-0.218	-0.001	-0.011	0.137	0.063
Px ₂ C ₄	-0.075	-0.248	-0.013	-0.004	-0.022	0.079	-0.075	-0.248	-0.013	-0.004	-0.022	0.079	-0.075	-0.248	-0.013
Px ₃ C ₄	0.066	0.344	0.047	0.052	0.038	0.101	0.066	0.344	0.047	0.052	0.038	0.101	0.066	0.344	0.047
sC ₅	-0.024	-0.014	0.020	-0.012	0.003	0.032	-0.024	-0.014	0.020	-0.012	0.003	0.032	-0.024	-0.014	0.020
Px ₁ C ₅	0.033	-0.074	-0.066	0.126	0.201	-0.036	0.033	-0.074	-0.066	0.126	0.201	-0.036	0.033	-0.074	-0.066
Px ₂ C ₅	0.106	0.074	-0.101	-0.170	0.092	-0.203	0.106	0.074	-0.101	-0.170	0.092	-0.203	0.106	0.074	-0.101
Px ₃ C ₅	-0.020	-0.058	-0.100	0.083	-0.107	-0.211	-0.020	-0.058	-0.100	0.083	-0.107	-0.211	-0.020	-0.058	-0.100
sC ₆	0.018	0.027	0.005	0.010	0.037	0.001	0.018	0.027	0.005	0.010	0.037	0.001	0.018	0.027	0.005
Px ₁ C ₆	-0.014	0.088	0.001	-0.136	-0.171	0.108	-0.014	0.088	0.001	-0.136	-0.171	0.108	-0.014	0.088	0.001
Px ₂ C ₆	-0.057	-0.076	0.030	0.152	-0.083	0.123	-0.057	-0.076	0.030	0.152	-0.083	0.123	-0.057	-0.076	0.030
Px ₃ C ₆	0.026	0.070	0.010	-0.096	0.089	0.151	0.026	0.070	0.010	-0.096	0.089	0.151	0.026	0.070	0.010
sC ₇	-0.024	-0.006	0.011	0.021	0.023	0.021	-0.024	-0.006	0.011	0.021	0.023	0.021	-0.024	-0.006	0.011
Px ₁ C ₇	0.044	-0.069	-0.023	0.106	0.183	-0.106	0.044	-0.069	-0.023	0.106	0.183	-0.106	0.044	-0.069	-0.023
Px ₂ C ₇	0.096	0.002	-0.060	-0.132	0.015	-0.131	0.096	0.002	-0.060	-0.132	0.015	-0.131	0.096	0.002	-0.060
Px ₃ C ₇	-0.014	-0.009	0.005	0.026	0.006	-0.085	-0.014	-0.009	0.005	0.026	0.006	-0.085	-0.014	-0.009	0.005
sC ₈	0.020	-0.006	0.008	0.020	0.006	0.024	0.020	-0.006	0.008	0.020	0.006	0.024	0.020	-0.006	0.008
Px ₁ C ₈	-0.076	0.043	-0.082	-0.141	-0.201	0.109	-0.076	0.043	-0.082	-0.141	-0.201	0.109	-0.076	0.043	-0.082
Px ₂ C ₈	0.080	-0.003	0.086	0.202	0.013	0.127	0.080	-0.003	0.086	0.202	0.013	0.127	0.080	-0.003	0.086
Px ₃ C ₈	0.037	-0.001	0.006	-0.158	0.129	0.093	0.037	-0.001	0.006	-0.158	0.129	0.093	0.037	-0.001	0.006
sC ₉	-0.043	0.011	0.050	-0.010	0.080	-0.018	-0.043	0.011	0.050	-0.010	0.080	-0.018	-0.043	0.011	0.050
Px ₁ C ₉	0.137	-0.039	-0.140	0.273	0.193	0.028	0.137	-0.039	-0.140	0.273	0.193	0.028	0.137	-0.039	-0.140
Px ₂ C ₉	0.205	-0.023	-0.237	-0.167	-0.105	0.011	0.205	-0.023	-0.237	-0.167	-0.105	0.011	0.205	-0.023	-0.237
Px ₃ C ₉	-0.096	-0.009	-0.019	0.125	-0.167	-0.059	-0.096	-0.009	-0.019	0.125	-0.167	-0.059	-0.096	-0.009	-0.019
sC ₁₀	0.015	-0.001	-0.005	0.007	0.008	0.034	0.015	-0.001	-0.005	0.007	0.008	0.034	0.015	-0.001	-0.005
Px ₁ C ₁₀	-0.092	0.034	0.074	-0.246	-0.222	0.045	-0.092	0.034	0.074	-0.246	-0.222	0.045	-0.092	0.034	0.074
Px ₂ C ₁₀	0.076	-0.063	0.083	0.209	-0.078	0.241	0.076	-0.063	0.083	0.209	-0.078	0.241	0.076	-0.063	0.083
Px ₃ C ₁₀	0.026	0.037	-0.002	-0.191	0.228	0.223	0.026	0.037	-0.002	-0.191	0.228	0.223	0.026	0.037	-0.002
sC ₁₁	0.014	-0.013	0.034	0.021	-0.020	0.039	0.014	-0.013	0.034	0.021	-0.020	0.039	0.014	-0.013	0.034
Px ₁ C ₁₁	-0.188	0.042	0.128	-0.026	0.001	-0.1010	-0.188	0.042	0.128	-0.026	0.001	-0.1010	-0.188	0.042	0.128
Px ₂ C ₁₁	-0.061	0.005	0.092	0.034	-0.021	0.054	-0.061	0.005	0.092	0.034	-0.021	0.054	-0.061	0.005	0.092
Px ₃ C ₁₁	0.017	0.001	-0.017	-0.027	0.045	0.083	0.017	0.001	-0.017	-0.027	0.045	0.083	0.017	0.001	-0.017
sC ₁₂	0.050	-0.004	-0.081	0.001	0.016	-0.003	0.050	-0.004	-0.081	0.001	0.016	-0.003	0.050	-0.004	-0.081
Px ₁ C ₁₂	0.256	-0.051	-0.233	0.069	-0.041	0.085	0.256	-0.051	-0.233	0.069	-0.041	0.085	0.256	-0.051	-0.233
Px ₂ C ₁₂	-0.054	0.012	-0.019	-0.054	0.127	-0.164	-0.054	0.012	-0.019	-0.054	0.127	-0.164	-0.054	0.012	-0.019
Px ₃ C ₁₂	-0.025	0.001	0.055	0.026	-0.052	0.035	-0.025	0.001	0.055	0.026	-0.052	0.035	-0.025	0.001	0.055
sC ₁₃	0.029	-0.017	0.073	-0.008	-0.023	0.016	0.029	-0.017	0.073	-0.008	-0.023	0.016	0.029	-0.017	0.073
Px ₁ C ₁₃	-0.294	0.120	-0.232	0.016	0.058	0.088	-0.294	0.120	-0.232	0.016	0.058	0.088	-0.294	0.120	-0.232
Px ₂ C ₁₃	0.039	-0.014	0.034	0.066	-0.164	0.250	0.039	-0.014	0.034	0.066	-0.164	0.250	0.039	-0.014	0.034
Px ₃ C ₁₃	0.050	-0.028	0.104	-0.043	0.068	-0.040	0.050	-0.028	0.104	-0.043	0.068	-0.040	0.050	-0.028	0.104
sC ₁₄	0.067	-0.040	0.025	0.145	0.230	-0.089	0.067	-0.040	0.025	0.145	0.230	-0.089	0.067	-0.040	0.025
Px ₁ C ₁₄	0.093	-0.033	0.044	-0.123	0.084	-0.213	0.093	-0.033	0.044	-0.123	0.084	-0.213	0.093	-0.033	0.044

Table II (contd.)

PC ₁₅	-0.008	-0.025	-0.015	-0.001	-0.040	0.016	0.003	-0.035	-0.020	-0.010	0.028	0.018	-0.035	0.018	-0.037	0.052	-0.120	-0.010
PC ₁₅	0.051	-0.076	-0.084	0.042	-0.118	0.093	0.050	0.075	-0.130	-0.012	0.082	-0.002	0.028	-0.002	0.007	-0.026	0.015	-0.017
PC ₁₅	0.065	0.088	0.062	-0.042	0.160	0.069	0.068	-0.014	0.135	0.082	-0.021	-0.158	0.067	-0.019	-0.053	-0.130	-0.026	-0.018
PC ₁₅	0.073	-0.020	-0.062	-0.032	-0.038	-0.133	0.012	0.005	-0.083	-0.048	-0.084	0.058	-0.050	0.015	-0.044	0.123	-0.037	0.108
PC ₁₆	0.073	0.038	0.030	-0.001	-0.028	-0.005	0.074	-0.004	0.025	0.030	-0.009	0.025	0.013	-0.004	0.012	-0.059	0.066	-0.013
PC ₁₆	0.006	-0.001	0.083	-0.001	-0.011	-0.038	0.012	-0.004	-0.008	0.113	-0.015	0.024	-0.053	0.021	-0.048	0.009	0.034	-0.021
PC ₁₆	-0.341	0.055	-0.064	0.010	0.070	-0.010	0.034	0.018	-0.073	-0.051	0.027	0.052	0.245	-0.104	0.244	-0.001	-0.062	0.032
PC ₁₆	-0.124	-0.014	0.024	0.010	-0.009	-0.029	-0.060	-0.005	-0.036	0.015	0.017	0.012	-0.067	0.129	0.068	0.004	-0.004	-0.001
PC ₁₆	-0.020	-0.001	0.033	-0.007	-0.002	-0.005	-0.192	0.008	0.002	0.042	-0.013	0.001	0.024	-0.007	-0.021	0.024	0.023	-0.013
PC ₁₇	-0.190	0.007	0.009	-0.001	-0.004	-0.008	0.017	-0.008	0.002	0.024	-0.006	0.006	0.024	-0.005	0.071	0.014	0.015	-0.003
PC ₁₇	0.192	-0.018	0.041	0.001	-0.017	-0.014	0.194	-0.011	-0.022	0.040	-0.005	0.014	0.145	-0.055	0.071	0.014	0.015	-0.003
PC ₁₇	-0.040	0.005	-0.012	0.007	0.012	-0.017	-0.023	0.002	0.005	0.034	-0.014	0.016	0.109	0.060	-0.127	0.014	0.028	-0.013
PC ₁₈	0.025	-0.001	0.005	0.011	0.008	-0.027	0.018	-0.001	0.001	0.008	0.017	-0.007	0.014	-0.006	0.015	-0.005	-0.009	0.002
PC ₁₈	0.071	0.003	-0.014	-0.020	-0.015	0.013	0.071	-0.001	-0.062	0.017	0.015	0.108	-0.018	0.008	-0.017	0.001	-0.001	-0.007
PC ₁₈	0.042	-0.036	0.016	0.056	-0.074	-0.077	0.037	-0.001	0.001	0.015	0.027	0.016	0.041	-0.019	0.054	-0.011	0.012	-0.015
PC ₁₈	0.082	0.014	0.018	0.070	0.090	-0.151	0.085	-0.006	0.026	0.039	0.119	0.080	-0.028	0.012	-0.026	-0.017	0.040	-0.089
PC ₁₉	0.002	0.002	-0.002	0.046	0.009	0.033	-0.009	-0.009	-0.015	0.105	-0.045	-0.016	0.038	-0.004	0.001	0.008	-0.005	0.006
PC ₁₉	0.004	-0.068	0.007	0.115	0.088	0.086	0.009	0.008	0.020	-0.001	0.093	-0.146	0.001	0.007	0.007	0.028	0.093	0.074
PC ₁₉	0.010	0.046	0.011	0.269	0.016	0.135	0.008	-0.009	0.020	0.025	0.229	0.041	0.001	-0.001	-0.003	-0.003	-0.012	-0.020
PC ₁₉	-0.002	-0.002	0.004	0.001	-0.005	-0.002	0.001	-0.007	0.001	-0.007	0.001	0.001	0.001	0.001	0.001	0.001	-0.002	-0.001
PC ₁₉	0.442	-0.035	0.083	-0.026	-0.014	-0.004	-0.001	-0.009	-0.048	0.003	-0.009	-0.147	0.001	0.001	0.001	0.001	-0.002	-0.001
PC ₁₉	0.497	0.068	-0.114	0.033	0.078	-0.018	-0.001	-0.129	-0.126	0.009	-0.147	0.001	0.001	0.001	0.001	0.001	-0.002	-0.001
PC ₁₉	0.116	-0.023	0.050	0.038	-0.015	-0.059	0.004	0.004	0.005	0.056	0.006	0.006	0.001	0.001	0.001	0.001	0.001	0.001
PC ₁₉	0.005	0.032	-0.005	0.146	0.015	0.146	-0.026	-0.026	0.073	-0.039	-0.007	0.007	0.001	0.001	0.001	0.001	0.001	0.001
PC ₁₉	0.003	0.021	-0.019	0.094	0.008	0.160	0.451	-0.029	0.086	-0.104	0.059	-0.060	0.001	0.001	0.001	0.001	0.001	0.001
PC ₁₉	0.003	0.009	-0.020	0.079	0.020	0.181	-0.502	-0.029	0.036	0.026	0.009	-0.009	0.001	0.001	0.001	0.001	0.001	0.001
PC ₁₉	0.006	0.001	0.117	0.043	0.223	0.043	0.073	-0.006	0.020	0.049	0.026	0.069	0.001	0.001	0.001	0.001	0.001	0.001
PC ₁₉	0.008	-0.001	-0.014	0.207	0.053	0.259	0.005	0.020	0.033	0.008	0.104	0.063	0.001	0.001	0.001	0.001	0.001	0.001
PC ₁₉	0.003	0.008	0.017	0.147	0.006	0.106	0.006	0.136	0.081	-0.033	-0.011	0.120	0.001	0.001	0.001	0.001	0.001	0.001
PC ₁₉	0.004	0.005	0.011	0.156	0.001	-0.057	0.006	0.072	0.037	-0.005	0.050	0.069	0.001	0.001	0.001	0.001	0.001	0.001
PC ₁₉	0.018	0.026	0.012	0.251	0.049	-0.119	0.000	0.106	-0.045	-0.092	-0.025	-0.095	0.001	0.001	0.001	0.001	0.001	0.001
PC ₁₉	0.016	-0.002	-0.004	0.004	0.097	0.084	0.004	-0.027	0.049	0.064	0.064	-0.087	0.001	0.001	0.001	0.001	0.001	0.001
PC ₁₉	0.030	-0.041	-0.032	0.179	0.034	-0.113	0.003	-0.026	0.026	-0.010	0.140	0.001	0.001	0.001	0.001	0.001	0.001	0.001
PC ₁₉	0.048	-0.006	-0.004	0.096	0.051	-0.121	0.003	0.061	0.082	-0.082	-0.010	0.140	0.001	0.001	0.001	0.001	0.001	0.001
PC ₁₉	0.048	0.037	0.035	0.008	0.138	0.088	-0.001	-0.019	0.021	0.117	-0.021	0.140	0.001	0.001	0.001	0.001	0.001	0.001
PC ₁₉	0.032	0.028	0.002	0.050	0.069	-0.102	0.003	-0.002	0.038	0.164	0.054	0.001	0.001	0.001	0.001	0.001	0.001	0.001
PC ₁₉	-0.071	-0.004	-0.052	-0.009	-0.012	0.047	0.003	0.028	0.041	0.303	0.057	0.001	0.001	0.001	0.001	0.001	0.001	0.001
PC ₁₉	0.081	0.042	0.029	-0.062	0.080	0.073	0.004	-0.002	0.023	0.315	-0.007	0.001	0.001	0.001	0.001	0.001	0.001	0.001
PC ₁₉	0.084	0.010	0.039	0.014	0.019	-0.015	0.003	-0.045	-0.003	0.213	-0.069	0.001	0.001	0.001	0.001	0.001	0.001	0.001
PC ₁₉	0.004	-0.049	-0.085	-0.167	-0.119	-0.087	0.004	-0.002	0.005	0.007	0.128	-0.069	0.001	0.001	0.001	0.001	0.001	0.001
PC ₁₉	0.007	-0.138	-0.121	-0.087	0.123	-0.103	0.003	-0.052	0.032	0.207	0.002	0.001	0.001	0.001	0.001	0.001	0.001	0.001
PC ₁₉	-0.015	0.195	-0.063	0.026	-0.002	-0.002	0.003	-0.002	0.043	0.012	0.090	-0.021	0.001	0.001	0.001	0.001	0.001	0.001
PC ₁₉	0.011	0.154	-0.068	-0.016	-0.191	-0.124	0.006	-0.007	-0.070	-0.012	-0.007	-0.007	0.001	0.001	0.001	0.001	0.001	0.001
PC ₁₉	0.033	0.179	-0.067	-0.097	-0.019	0.007	0.085	-0.010	0.069	0.030	-0.069	-0.110	0.001	0.001	0.001	0.001	0.001	0.001
PC ₁₉	-0.001	-0.154	-0.138	-0.049	-0.027	-0.023	0.006	0.006	0.018	-0.041	0.029	-0.007	0.001	0.001	0.001	0.001	0.001	0.001
PC ₁₉	0.037	-0.138	-0.024	-0.071	-0.173	0.123	0.028	0.181	-0.178	-0.070	-0.070	0.102	0.001	0.001	0.001	0.001	0.001	0.001
PC ₁₉	0.029	0.136	-0.079	-0.008	-0.135	0.093	0.028	-0.025	-0.139	-0.054	-0.031	-0.060	0.001	0.001	0.001	0.001	0.001	0.001
PC ₁₉	0.016	0.008	-0.021	-0.029	-0.011	0.031	0.003	-0.009	-0.088	-0.088	-0.100	0.001	0.001	0.001	0.001	0.001	0.001	0.001
PC ₁₉	0.018	0.045	0.007	-0.003	-0.118	0.050	0.003	0.004	0.071	-0.005	-0.026	-0.125	0.001	0.001	0.001	0.001	0.001	0.001
PC ₁₉	0.004	0.101	-0.023	-0.045	-0.097	-0.044	0.001	0.006	0.004	-0.022	-0.038	0.014	0.001	0.001	0.001	0.001	0.001	0.001
PC ₁₉	-0.019	-0.026	0.110	-0.076	0.062	-0.059	0.002	-0.073	-0.006	-0.033	0.127	0.001	0.001	0.001	0.001	0.001	0.001	0.001
PC ₁₉	0.005	-0.070	-0.078	-0.096	0.034	-0.022	0.003	-0.031	-0.031	-0.040	0.169	-0.048	0.001	0.001	0.001	0.001	0.001	0.001
PC ₁₉	0.003	-0.070	-0.078	-0.096	0.034	-0.022	0.003	-0.031	-0.031	-0.040	0.169	-0.048	0.001	0.001	0.001	0.001	0.001	0.001
PC ₁₉	0.003	-0.070	-0.078	-0.096	0.034	-0.022	0.003	-0.031	-0.031	-0.040	0.169	-0.048	0.001	0.001	0.001	0.001	0.001	0.001
PC ₁₉	0.003	-0.070	-0.078	-0.096	0.034	-0.022	0.003	-0.031	-0.031	-0.040	0.169	-0.048	0.001	0.001	0.001	0.001	0.001	0.001
PC ₁₉	0.003	-0.070	-0.078	-0.096	0.034	-0.022	0.003	-0.031	-0.031	-0.040	0.169	-0.048	0.001	0.001	0.001	0.001	0.001	0.001
PC ₁₉	0.003	-0.070	-0.078	-0.096	0.034	-0.022	0.003	-0.031	-0.031	-0.040	0.169	-0.048	0.001	0.001	0.001	0.001	0.001	0.001
PC ₁₉	0.003	-0.070	-0.078	-0.096	0.034	-0.022	0.003	-0.031	-0.031	-0.040	0.169	-0.048	0.001	0.001	0.001	0.001	0.001	0.001
PC ₁₉	0.003	-0.070	-0.078	-0.096	0.034	-0.022	0.003	-0.031	-0.031	-0.040	0.169	-0.048	0.001	0.001	0.001	0.001	0.001	0.001
PC ₁₉	0.003	-0.070	-0.078	-0.096	0.034	-0.022	0.003	-0.031	-0.031	-0.040	0.169	-0.048	0.001	0.001	0.001	0.001	0.001	0.001
PC ₁₉	0.003	-0.070	-0.078	-0.096	0.034	-0.022	0.003	-0.031	-0.031	-0.040	0.169	-0.048	0.001	0.001	0.001	0.001	0.001	0.001
PC ₁₉	0.003	-0.070	-0.078	-0.096	0.034	-0.022	0.003	-0.031	-0.031	-0.040	0.169	-0.048	0.001	0.001	0.001			

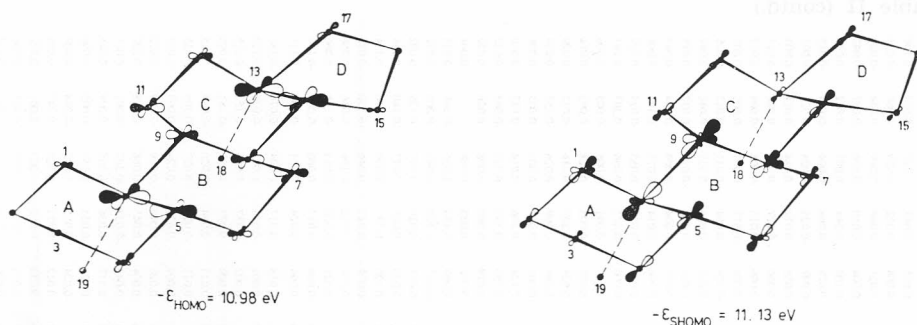


Figure 1. The transversal (T-type) and longitudinal (L-type) ribbon-like MO's of 5 α -androstane, according to MNDO results.

nounced shift of σ -systems towards higher energies, whereas in the low energy region a new system appears. It corresponds to the carbonyl oxygen lone pair ionization. Nevertheless it is not strictly localized on the oxygen atom but is rather mixed with the steroid ribbon-type orbital giving rise to significant delocalization. This is supported not only by MNDO calculations (see Table II) but is also evident from the vibrational structure broadening in the related PE band.

The details of the perturbation caused by this highly electronegative group can be closely studied in the PE spectra of a series of monoketones over diketones to triketones. According to the calculated electron density distribution, the carbonyl group would mix either with T- or L-type orbital depending on the position of substitution. Namely, in 3-one case HOMO is of T-type, in 17-one case it is of L-type, while in 11-one case HOMO could be described as a mixture of both types of orbitals (T- plus L-type). In 3,17-dione HOMO and SHOMO are closely related to HOMO's of the corresponding monoketones, *i. e.*, it seems that no interaction has taken place. It should be pointed out that no geometry optimization has been performed for 4, since it is known that through-bond interaction is sensitive not only to the number and parity of bonds between two centers, but also on the assumed molecular geometry.

The interpretation of the PE spectrum of 3,11,17-trione could be supported by the results available from monoketones and 3,17-diketone. The low energy part of the spectrum of 6 reveals three electron systems. *Per analogiam* with PE spectra of 2 and 5 one deduces that the middle system appearing at 9.09 eV could be assigned to the almost unperturbed carbonyl oxygen lone pair in A ring. On the contrary, the other two systems at 8.90 and 9.50 eV, respectively, should correspond to ionizations from significantly mixed wave functions of 11- and 17-carbonyl groups. That is nicely reproduced in diagrammatic representation of MNDO electron density distribution (Figure 2). Even the experimentally observed splitting amounting to 0.60 eV is correctly predicted by this calculation ($\Delta\epsilon^{\text{MNDO}}(n_0) = 0.59$ eV). Taking into account relative orien-

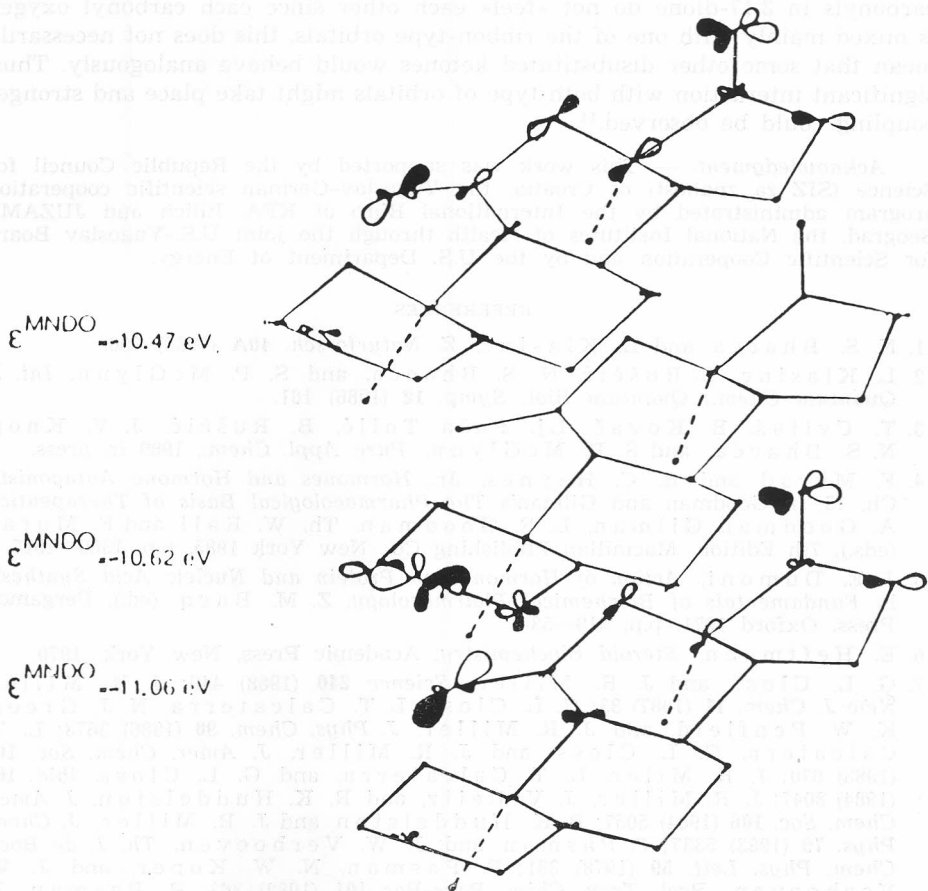


Figure 2. The three highest occupied MO's of 5 α -androstane-3,11,17-trione on the basis of MNDO calculations.

tation of two oxygen atoms and their interatomic distance which, according to the MNDO calculations amounts to 540 pm, this prominent coupling could be due to through bond interaction. On the other hand, no interaction is observed between the carbonyl group in ring A and either of the carbonyl groups in C or D rings. This is confirmed by MNDO calculations, since no participation from oxygen lone pair electrons in 11- and 17-positions is found in the SHOMO of 6 (which corresponds to the oxygen lone pair on 3-position; see Figure 2). Such a theoretical result is not surprising as these two oxygen atoms are 8 and 10 σ_{CC} bonds apart from the one in 3-position, respectively.

The close study of long-range interactions gives insight into the type and extent of orbital mixing. As we have seen the position of a carbonyl group is by no means irrelevant (*e.g.*, positions 3- vs. 17- vs. 11-). In each case the type of orbital mixing is different and therefore the effect of coupling between two or three carbonyl groups differs significantly. Although two

carbonyls in 3,17-dione do not »feel« each other since each carbonyl oxygen is mixed mainly with one of the ribbon-type orbitals, this does not necessarily mean that some other disubstituted ketones would behave analogously. Thus, significant interaction with both type of orbitals might take place and stronger coupling could be observed.¹⁴

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SAŽETAK**Fotoelektronski spektri, MNDO izračunavanja i elektronska struktura nekih zasićenih steroida**

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Fotoelektronski spektri, semiempirijska MNDO SCF MO izračunavanja (uz pretpostavku valjanosti Koopmanova teorema) te empirijske činjenice korištene su za određivanje strukture (valentnih) elektrona 5 α -androstana (1), 5 α -androstan-3-ona (2), 5 α -androstan-11-ona (3), 5 α -androstan-17-ona (4), 5 α -androstan-3,17-diona (5) i 5 α -androstan-3,11,17-triona (6).